**Conjugate Acid-Base Pairs**

According to the Bronsted-Lowry Theory of Acids and Bases, after an acid gives up its proton, it is capable of getting back that proton and acting as a base. The conjugate base is what is left after the acid gives up a proton.

*Fill in the blanks in the table with the appropriate conjugate.*

|  |  |  |
| --- | --- | --- |
| **Conjugate Pairs** | | |
|  | **ACID** | **BASE** |
| 1. | H2SO4 |  |
| 2. | H3PO4 |  |
| 3. |  | F -1 |
| 4. |  | NO3 -1 |
| 5. | H2PO4 -1 |  |
| 6. | H2O |  |
| 7. |  | SO4 -2 |
| 8. | HPO4 -2 |  |
| 9. | NH4 +1 |  |
| 10. |  | H2O |
| 11. |  | H2PO3 -1 |
| 12. |  | HSO3 -1 |